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**Raymond, Ashton** [US/US]; 1574 South 1175 East,  
Bountiful, UT 84010 (US).

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(74) Agent: **BURFORD, Anthony, F.**; W.H. Beck, Greener &  
Co., 7 Stone Buildings, Lincoln's Inn, London WC2A 3SZ  
(GB).

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(71) Applicant (*for all designated States except US*): **AIR  
PRODUCTS AND CHEMICALS, INC.** [US/US]; 7201  
Hamilton Boulevard, Allentown, PA 18195-1501 (US).

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(72) Inventor: **DYER, Paul, Nigel** (deceased).

(72) Inventors; and

(75) Inventors/Applicants (*for US only*): **CAROLAN,**  
**Michael, Francis** [US/US]; 1091 Treeline Drive, Al-  
lentown, PA 18103 (US). **BUTT, Darryl, P.** [US/US];  
8203 SW 16th Place, Gainesville, FL 32607 (US). **VAN**  
**DOORN, Rene, Hendrik, Elias** [NL/DE]; Dr.-Loef-  
felladstrasse 16, 86609 Donauworth (DE). **CUTLER,**

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(54) Title: MIXED CONDUCTING MEMBRANES FOR SYNGAS PRODUCTION

(57) Abstract: Multicomponent metallic oxides, which are particularly suited toward use in fabricating mixed conducting mem-  
branes used in processes for producing syngas, are novel non-stoichiometric, A-site rich compositions represented by the formula  
(Ln<sub>x</sub>Ca<sub>1-x</sub>)<sub>y</sub>FeO<sub>3-δ</sub>, wherein Ln is La or a mixture of lanthanides comprising La, 1.0>x>0.5, 1.1≥y>1.0 and δ is a number which  
renders the metallic oxide charge neutral.

## Mixed Conducting Membranes for Syngas Production

This invention was made with United States Government support under Cooperative Agreement DE-FC26-97PC96052 awarded by the United States Department of Energy. The Government has certain rights in this invention.

Synthesis gas (syngas) containing hydrogen and carbon oxides is an important feedstock for the production of a wide range of chemical products. Syngas mixtures having controlled ratios of hydrogen and carbon monoxide are catalytically reacted to produce liquid hydrocarbons and oxygenated organic compounds including methanol, acetic acid, dimethylether, oxoalcohols and isocyanates. The syngas product can be further processed and separated to yield high purity hydrogen and carbon monoxide. The cost of generating the syngas is frequently the largest part of the total cost of preparing these products.

Two major reaction routes are commonly used by industry for syngas production, namely steam reforming of light hydrocarbons, primarily natural gas, naphtha and refinery offgases, and the partial oxidation of carbon-containing feed stocks ranging from natural gas to high molecular weight liquid or solid carbonaceous materials. Autothermal reforming is an alternate process which uses a light hydrocarbon feed which combines features of partial oxidation and steam reforming reactions in a single reactor. A concise review of such processes is described in US-A-6,077,323. Such processes typically require oxygen in purities of greater than 95 vol%, which is available from cryogenic air separation in large tonnage volumes or pressure swing absorption for smaller volumes.

Alternative processes have been developed for syngas production wherein oxygen necessary to conduct the partial oxidation reaction is provided *in situ* by the separation of air at high temperatures using solid-state membranes which conduct oxygen ions and electrons under operating conditions. Solid-state membranes which conduct oxygen ions and electrons are also known as mixed conducting membranes. Such mixed conducting membranes can be used in combination with appropriate catalysts to produce syngas in a membrane reactor eliminating the need for a separate oxygen production step. A membrane reactor typically has one or more reaction zones, wherein each reaction zone comprises a mixed conducting

membrane which separates each reaction zone into an oxidant side and a reactant side.

Multicomponent metallic oxides are represented in the art by formulae which present one or more "A-site" metals and one or more "B-site" metals. By way of example, US-A-5,306,411 discloses certain multicomponent metallic oxides having the perovskite structure represented by the formula  $A_s A'_t B_u B'_v B''_w O_x$ , wherein A represents a lanthanide, Y or a mixture thereof; A' represents an alkaline earth metal or mixture thereof; B represents Fe; B' represents Cr, Ti or a mixture thereof; and B'' represents Mn, Co, V, Ni, Cu or a mixture thereof, and s, t, u, v, w and x each represent a number such that s/t equals from about 0.01 to about 100; u equals from about 0.01 to about 1; v equals from about 0.01 to about 1; w equals from 0 to about 1; x equals a number that satisfies the valences of A, A', B, B' and B'' in the formula; provided that  $0.9 < (s+t)/(u+v+w) < 1.1$ . In a preferred embodiment A' is calcium or strontium and B'' represents Mn or Co or a mixture thereof. These multicomponent metallic oxides require chromium or titanium as a B-site element.

Multicomponent metallic oxides depicted by formulae presenting A-site metals and B-site metals may be stoichiometric compositions, A-site rich compositions or B-site rich compositions. Stoichiometric compositions are defined as materials wherein the sum of the A-site metal stoichiometric coefficients equals the sum of the B-site metal stoichiometric coefficients. A-site rich compositions are defined as materials wherein the sum of the A-site metal stoichiometric coefficients is greater than the sum of the B-site metal stoichiometric coefficients. B-site rich compositions are defined as materials wherein the sum of the B-site metal stoichiometric coefficients is greater than the sum of the A-site metal stoichiometric coefficients.

US-A-6,033,632 discloses a solid-state membrane for use in a catalytic membrane reactor which utilizes a membrane fabricated from a multicomponent metallic oxide having the stoichiometry  $A_{2-x} A'_x B_{2-y} B'_y O_{5+z}$ , wherein A is an alkaline earth metal ion or mixture of alkaline earth metal ions; A' is a metal ion or mixture of metal ions where the metal is selected from metals of the lanthanide series and yttrium; B is a metal ion or mixture of metal ions, wherein the metal is selected from the 3d transition metals and the Group 13 metals; B' is a metal ion or mixture of metal ions where the metal is selected from the 3d transition metals, the Group 13

metals, the lanthanides and yttrium; x and y are independently of each other numbers equal to or greater than zero and less than 2; and z is a number that renders the compound charge neutral. In a preferred embodiment the 3d transition metal is Fe and the Group 13 metal is Ga, whereas A' preferably is La and A is Sr and Ba.

US-A 5,356,728 and US-A 5,580,497 disclose cross-flow electrochemical reactor cells formed from multicomponent metallic oxides which demonstrate electron conductivity and oxygen ion conductivity at elevated temperatures.

10 According to both documents, suitable multicomponent metallic oxides are represented by  $(\text{Sr}_{1-y}\text{M}_y)_\alpha(\text{Fe}_{1-x}\text{Co}_x)_{\alpha+\beta}\text{O}_\delta$ , wherein M is a metal selected from elements having atomic number in a range from 56 to 71, calcium and yttrium, x is a number in a range from 0.01 to 0.95, y is a number in a range from 0.01 to 0.95,  $\alpha$  is a number in a range from 1 to 4,  $\beta$  is a number in a range upward from 0 to 20, such

15 that  $1.1 < (\alpha + \beta) / \alpha \leq 6$ , and  $\delta$  is a number which renders the compound charge neutral.

US-A 6,056,807 teaches a fluid separation device capable of separating oxygen from an oxygen-containing gaseous mixture which utilizes at least one solid-

20 state membrane comprising a dense mixed conducting multicomponent metallic oxide layer formed from a metallic oxide represented by the formula  $\text{Ln}_x\text{A}'_x\text{A}''_x\text{B}_y\text{B}'_y\text{O}_{3-z}$ , wherein Ln is an element selected from the f block lanthanides, A' is selected from Group 2, A'' is selected from Groups 1, 2 and 3 and the f block lanthanides and B and B' are independently selected from the d block transition

25 metals, excluding titanium and chromium, wherein  $0 \leq x < 1$ ,  $0 < x' \leq 1$ ,  $0 \leq x'' < 1$ ,  $0 < y < 1.1$ ,  $0 \leq y' < 1.1$ ,  $x + x' + x'' = 1.0$ ,  $1.1 > y + y' > 1.0$  and z is a number which renders the compound charge neutral. This reference discloses B-site rich non-stoichiometric compositions because the sum of the x indices is 1.0 and the sum of the y indices is greater than 1.0.

30 US-A 5,712,220 presents a class of multicomponent metallic oxides which are well suited toward use in fabricating components used in solid-state oxygen separation devices. While the reference relates primarily to B-site rich compositions, the reference discloses A-site rich non-stoichiometric compositions represented by

35 the formula  $\text{Ln}_x\text{A}'_x\text{A}''_x\text{B}_y\text{B}'_y\text{O}_{3-z}$  wherein Ln is an element selected from the f

block lanthanides, A' is selected from Group 2, A'' is selected from Groups 1, 2 and 3 and the f block lanthanides and B, B' and B'' are independently selected from the d block transition metals, excluding titanium and chromium, wherein  $0 \leq x < 1$ ,  $0 < x' < 1$ ,  $0 \leq x'' < 1$ ,  $0 < y < 1.1$ ,  $0 < y' < 1.1$ ,  $0 < y'' < 1.1$ ,  $x + x' + x'' = 1.0$ ,  $0.9 < y + y' + y'' < 1.0$  and z is a number which renders the compound charge neutral wherein such elements are represented according to the Periodic Table of the Elements adopted by IUPAC.

A solid-state membrane employed in a process for making syngas is exposed to severe reaction conditions such as temperatures above 600°C, a large pressure difference across the solid-state membrane, a highly oxidizing environment on one surface, and a water, hydrogen, methane, carbon monoxide and carbon dioxide containing reactant gas stream on the other surface. Therefore, the solid-state membrane must have sufficiently high oxygen flux, a sufficient chemical stability in the syngas and air environments, a sufficiently low creep rate under the applied mechanical load, a sufficient resistance to demixing of the metal cations and a sufficiently low chemical expansion under the membrane operating conditions.

Numerous compositions known in the art for fabricating solid-state membranes do not adequately meet all the above criteria. Although some compositions are known to meet the oxygen flux criteria, for example, these compositions typically may not meet one or more other criteria as listed above. These criteria are nevertheless highly sought after for development of an economically viable technology based on solid-state membranes.

Researchers continue to search for suitable solid-state membranes that will economically and reliably produce syngas through the oxidation of methane and partially reformed methane feed stocks. More in detail, researchers are searching for mixed conducting multicomponent metallic oxides suitable for use in fabricating the dense layer of a solid-state membrane which meet the above criteria.

The present Inventors have discovered a new class of A-site rich non-stoichiometric multicomponent metallic oxides which are particularly suited toward use in solid-state membranes suitable for use in processes for producing synthesis gas (syngas). These metallic oxides overcome problems associated with many prior art metallic oxides by providing a favourable balance of oxygen permeance, resistance to

degradation, favourable sintering properties and coefficients of thermal expansion which are compatible with other materials used to fabricate solid-state membranes.

The metallic oxides according to the invention are represented by Formula (I):



wherein

Ln is La or a mixture of lanthanides comprising La, and wherein

$$1.0 > x > 0.5$$

$$1.1 \geq y > 1.0 \text{ and}$$

10  $\delta$  is a number which renders the metallic oxide charge neutral.

In a preferred embodiment,  $0.98 > x > 0.75$  and  $1.05 \geq y \geq 1.01$ .

For purposes of interpreting Formula (I), the A-sites of the multicomponent  
15 metallic oxide are represented by Ln and Ca which bear stoichiometric coefficients x and 1-x, respectively. The B-site of the multicomponent metallic oxide according to the Formula is represented by Fe which bears the stoichiometric coefficient of 1 (not shown). The sum of the A-site stoichiometric coefficients ((x) + (1-x)) multiplied by y wherein  $1.0 > x > 0.5$  and  $1.1 > y > 1.0$  is always greater than the stoichiometric  
20 coefficient of the B-site, Fe, which is one. Thus, the metallic oxides according to this invention are A-site rich materials. These A-site rich metallic oxides have the perovskite structure and are particularly suited toward use as a dense layer in a solid-state membrane suitable for use in processes for making syngas.

25 The metallic oxides of this invention may be prepared by conventional ceramic methods wherein respective amounts of metal oxides, carbonates, hydroxides, nitrates, or oxalates are mixed to achieve the desired cation fraction of the metal cation thereof. Suitable mixing can be effected by conventional means such as a technique selected from attrition milling, vibratory milling, ball milling and high shear mixing. The resulting  
30 mixture is calcined to achieve a homogenous single phase, or nearly single phase, ceramic powder.

The metallic oxides according to Formula (I) can be used to form the dense layer of a solid-state membrane. Thus the invention also presents solid-state membranes  
35 which comprise a dense layer formed from a metallic oxide represented by Formula (I):



wherein

$\text{Ln}$  is La or a mixture of lanthanides comprising La, and wherein

$1.0 > x > 0.5$

$1.1 \geq y > 1.0$  and

5  $\delta$  is a number which renders the metallic oxide charge neutral.

The term "dense layer" used herein means a membrane layer which has substantially no connected through porosity so that the dense layer is substantially impervious to the passage of oxygen-containing or reactant gases. Minor amounts of  
10 transport of gases across the dense layer may occur without detriment to utility of the invention. The dense layer conducts oxygen anions and electronic species under operating conditions and in this sense are permeable to oxygen ions, but not molecular oxygen.

15 In a preferred embodiment, the dense layer of the solid-state membrane is formed from a metallic oxide according to the formula wherein  $0.98 > x > 0.75$  and  $1.05 \geq y \geq 1.01$ .

While this Application primarily addresses solid-state membranes comprising a  
20 dense layer formed from such compositions, additional layers of the solid-state membrane may also be fabricated including metallic oxides according to the Formula.

In particular, the solid state membranes may further comprise any number of additional layers to enhance performance and durability. Such additional layers may  
25 include a porous mixed conducting multicomponent metallic oxide layer contiguous to the dense layer. The term "porous" used herein means that the relevant layer possesses a network of pores such that the oxygen-containing feed gas can diffuse through the porous layer to make contact with the dense layer of the solid-state membrane.

30

The solid-state membranes which comprise a dense layer and any number of additional layers may be fabricated into a variety of shapes including flat plates or tubes. These solid-state membranes possess two exterior sides, referred to as the first side and the second side.

35

Such solid-state membranes may also include a catalyst on the first side, a catalyst on the second side or a catalyst on the first side and the second side, meaning the surfaces of the membrane which will be in contact with the oxygen-containing feed gas and the methane-containing reactant gas during operating of the syngas process of this invention.

By way of example, the first side may be referred to as the reactant side and the second side may be referred to as the oxidant side. Suitable catalysts to be deposited onto the reactant side of the solid-state membrane are conventional reforming catalysts or partial oxidation catalysts such as a metal or an oxide of a metal selected from Groups 5, 6, 7, 8, 9, 10, 11 of the Periodic Table of the Elements according to the International Union of Pure and Applied Chemistry. Preferred metals or oxides of metals are selected from nickel, cobalt, iron, platinum, palladium, and silver.

By way of example, the second side of the solid-state membrane may be referred to as the oxidant side. Suitable catalysts to be deposited onto the oxidant side are conventional oxygen reduction catalysts such as a metal or an oxide of a metal selected from the Groups 2, 5, 6, 7, 8, 9, 10, 11 and 15 and the F block lanthanides of the Periodic Table of the Elements according to the International Union of Pure and Applied Chemistry. Preferred metals or oxides of metals are selected from platinum, palladium, ruthenium, gold, silver, bismuth, barium, vanadium, molybdenum, cerium, praseodymium, cobalt, rhodium and manganese. Optionally, the catalyst may be any multicomponent metallic oxide which catalyzes the desired reaction.

The solid-state membrane of the present invention comprises a dense layer of a composition according to Formula (I), and optionally one or more additional layers including, but not limited to porous layers and/or catalytic layers. Porous layers contiguous to the dense layer of solid-state membranes of this invention are preferably formed from a mixed conducting multicomponent metallic oxide although each porous layer, including those layers contiguous to the dense layer, may be formed from a electrically-conducting multicomponent metallic oxide, an oxygen ionically-conductive multicomponent metallic oxide or an inert material which does not conduct electrons or oxygen ions under process operating conditions.

The solid-state membrane comprising its respective layers is fabricated to possess a thickness sufficient to be mechanically stable to withstand the stresses



associated with process operation, yet not so thick as to substantially limit the oxygen permeation rate through the solid-state membrane. Solid-state membranes can be fabricated in a variety of shapes appropriate for a particular reaction zone within a reactor, including disks, tubes, closed-end tubes, planar designs or as reactor cores for cross-flow operation. For example, one can envision the two sides of a coin in the case of a flat plate or the interior and exterior surfaces of a tube.

The solid-state membranes having a dense layer formed from the metallic oxides according to Formula (I) are particularly suited for use in processes for making syngas because such compositions have sufficiently high oxygen flux, a sufficient chemical stability in the syngas and air environments, a sufficiently low creep rate, a sufficient resistance to demixing of the metal cations and a sufficiently low chemical expansion under membrane operating conditions.

In the case where the solid-state membrane comprises a dense layer formed from the metallic oxides of this invention and one or more porous layers, a mixed conducting multicomponent metallic oxide layer is preferably employed for each porous layer contiguous to the dense layer. The porous mixed conducting oxide layer contiguous with the dense mixed conducting layer typically has a thickness ranging from 1 micrometer to 2 millimeters. Porous layers not in contact with the dense layer, whether or not formed from a mixed conducting oxide, can be as thick as desired to ensure mechanical strength to the extent that the ultimate thickness does not impede gas diffusion. Typically, the total thickness of the solid-state membrane is less than 5 mm, although solid-state membranes having a greater thickness are also contemplated.

The thickness of the dense mixed conducting layer typically ranges from 0.01 micrometer to 500 micrometers although preferably, the dense layer is fabricated as thinly as permitted by structural integrity considerations and has a thickness of less than 100 micrometers.

The thickness of each porous layer of the solid-state membrane, if so employed, can be varied to ensure sufficient mechanical strength of the solid-state membrane. The desired thickness of each porous layer is regulated according to the following considerations. First, the porosity and average pore radius of each porous layer should be regulated such that oxygen flux is not impeded while maintaining sufficient mechanical strength. Second, the pores or channels of pores within each porous layer

should be wide enough so that oxygen flux is not impeded, but not so wide as to cause pore filling during fabrication or failure of the dense layer of the solid-state membrane during operation. Third, each porous layer should be compatible with the dense layer in terms of chemical reactivity and adhesion and thermal expansion to reduce problems associated with cracking and delamination.

The solid-state membranes can be fabricated by depositing a dense layer of the metallic oxides according to Formula (I) in a desired thickness onto one or more porous layers by known techniques. For example, the solid-state membrane can be manufactured by first forming a porous mixed conducting multicomponent metallic oxide layer from relatively coarse sized particles of the desired metal oxides. A slurry of finer particles of a compatible metallic oxide according to Formula (I), optionally including conventional binders, plasticizers and/or solvents, may then be coated onto the porous body and cured to the green state to form a green body, the two layer system then being sintered to form the solid-state membrane. Alternatively, the slurry can be spray dried to obtain a powder. This powder can then be isostatically pressed to form a green body. The green body obtained by these conventional ceramic techniques is then sintered.

Alternatively, the solid-state membranes of the present invention can be prepared by applying a dense layer of the desired mixed conducting oxide according to Formula (I) onto a desired porous substrate by conventional chemical vapour deposition techniques followed by sintering to obtain the desired dense layer. In order to obtain an optimal dense coating, a smaller average pore radius in the surface of the porous layer may be used compared to the average pore radius in the bulk. This may be achieved by using two or more porous layers which differ in properties such as pore radius and porosity.

Green bodies can be sintered in electric, gas, microwave, plasma, or other furnaces. Batch or continuous furnaces can effectively be used. Sintering temperatures should typically be below 1600°C, preferably in the range of 1200 to 1400 °C in order to maximize furnace and setter life times.

The dense layer of the membrane is preferably sintered to a density of at least 95 % theoretical density, preferably at least 99 % theoretical density. Theoretical density may be measured according to known methods, see e.g. G.H. Stront, L.H.

Jensen, in "X-Ray Structure Determination, A Practical Guide", the Macmillan Company, Ltd., London, p. 80-81

The metallic oxide of the present invention and the solid-state membranes comprising the same may be stack fired, hang fired, or fired by use of any other means to minimize interaction of the solid-state membrane with the setter. Suitable setters known in the art may be used. In any case it is desired to prevent sticking between the setter and the solid-state membrane.

Catalysts can be deposited onto or incorporated within any desired layer of the solid-state membrane by employing conventional techniques. By way of example, a desired catalyst may be applied to the desired side of the solid-state membrane by techniques such as dip coating with a slurry or suspension of powder or spraying a suspension of powder, followed by drying and firing. Alternatively, the catalyst can be incorporated into a dense layer, porous layer or other layer by mixing the catalyst with the powder of the multicomponent metallic oxide which shall make up the layer prior to sintering the same to form a catalyst-containing layer of the solid-state membrane.

It is emphasized that the entire surface of the solid-state membrane does not have to be coated with catalyst in order to achieve the benefits afforded by use of the catalyst. For example, any selected pattern of catalyst may be deposited onto the surface by screen printing, masking and other techniques. Such patterns can be designed and applied according to currently used techniques which are well known in the art.

Solid-state membranes formed from the metallic oxides of this invention exhibit excellent long term oxygen flux stability and stability against demixing under operating conditions. Further, such solid-state membranes are stable under much higher partial pressures of carbon dioxide than the membrane compositions of the prior art. Finally, the solid-state membranes of this invention exhibit lower chemical expansion under operating conditions than solid-state membranes formed from prior art compositions.

The invention also presents a process for producing a synthesis gas product comprising hydrogen and carbon monoxide comprising the steps of:

i) providing a reaction zone having an oxidant side and a reactant side which are separated by a solid-state membrane comprising a dense layer formed from a metallic oxide represented by Formula (I)



wherein

Ln is La or a mixture of lanthanides comprising La, and wherein

$$1.0 > x > 0.5$$

$$1.1 \geq y > 1.0 \text{ and}$$

$\delta$  is a number which renders the metallic oxide charge neutral,

ii) contacting a heated oxygen-containing feed gas with the oxidant side of the said membrane at an oxidant feed temperature and an oxidant gas feed pressure;

iii) contacting a heated methane-containing reactant gas with the reactant side of the said membrane at a reactant gas feed temperature and a reactant gas feed pressure;

whereby oxygen from the oxidant side of the reaction zone permeates through the solid-state membrane to the reactant side of the reaction zone and reacts with the methane-containing reactant gas to form the synthesis gas product;

iv) withdrawing the synthesis gas product from the reactant side of the reaction zone; and

v) withdrawing an oxygen depleted gas stream from the oxidant side of the reaction zone.

The term "an oxygen-containing feed gas" includes gases and mixtures of gases in which at least one of the component gases is oxygen or an oxide. The oxygen or oxide component of the gas is capable of being reduced at the surface of the solid-state membrane of this invention. The oxygen-containing gas may include carbon, nitrogen, and sulfur oxides ( $\text{CO}_x$ ,  $\text{NO}_x$  and  $\text{SO}_x$ ) among others, and gas mixtures in which an oxide is a component, e.g.  $\text{NO}_x$  in an inert gas or in another gas not reactive with the solid-state membrane. The term also includes mixtures of oxygen in other gases, e.g.  $\text{O}_2$  in air.

The term "a methane-containing reactant gas" means gases or mixtures of gases containing methane, including but not limited to methane, natural gas (whose major component is methane), and gaseous hydrocarbons including light hydrocarbons (as this term is defined in the chemical arts). The term also refers to mixtures of methane which include inert gases, or mixtures of such components with oxygen-containing species, such as CO, CO<sub>2</sub> or H<sub>2</sub>O. Preferably, the methane-containing reactant gas contains water and the water to carbon molar ratio is between 0.5 to 5, the water to carbon molar ratio being defined as the number of molecules of water in the reactant gas divided by the total number of carbon atoms present in the hydrocarbons and carbon monoxide in the reactant gas.

The oxygen containing feed gas in step (ii) is preferably heated by direct combustion with a fuel in a direct-fired combustor to produce a hot, pressurized combustion product to provide the heated oxygen-containing feed gas.

15

Optionally, the process further comprises the steps of:

vi) contacting a heated gaseous stream comprising steam and one or more hydrocarbons with at least one catalyst which promotes steam reforming of hydrocarbons to form a partially reformed intermediate gas comprising at least methane, hydrogen and carbon oxides; and

20

vii) introducing the partially reformed intermediate gas into the reactant side of the reaction zone of step iii).

25

The term "partially reformed intermediate gas" as used herein refers to the product gas formed by the reaction of steam with a feed gas containing one or more hydrocarbons heavier than methane, and optionally containing ethane, wherein the reaction products comprise methane, carbon oxides, hydrogen, and steam (defined herein as vaporized or gaseous water). The partially reformed intermediate gas preferably is essentially free of hydrocarbons heavier than methane, which means that this gas contains less than 100 ppm by volume of hydrocarbons heavier than methane.

30

Process operating conditions can be readily varied depending upon the particular methane-containing reactant gas or oxygen-containing feed gas utilized.

35

Typically, the reactant feed gas temperature is between 950°F (510°C) and 1400°F (760°C) and the synthesis gas product outlet temperature from the reactant side is greater than 1500°F (815°C). The oxidant gas feed pressure preferably is between 1 psig (0.07 barG) and 45 psig (3 barG). The reactant gas feed pressure preferably is  
5 between 100 psig (7 barG) and 900 psig (62 barG) and the oxidant gas feed temperature is up to 200°F (95°C) greater than the reactant gas feed temperature. The oxidant gas feed temperature is preferably less than the oxygen-depleted oxidant gas outlet temperature.

10 The following is a description by way of example only, and with reference to the accompanying drawings, of specific embodiments of the invention. In the drawings: -

Fig. 1 is a process flow diagram of an embodiment of the process of the invention for producing a synthesis gas product;

15

Fig. 2 is a graph showing oxygen flux performance versus time for two solid-state membranes of the invention;

Fig. 3 is a graph showing equilibrium concentrations of CO<sub>2</sub> and O<sub>2</sub> for  
20 decomposition of La<sub>0.85</sub>Ca<sub>0.15</sub>FeO<sub>3-δ</sub> and La<sub>0.85</sub>Sr<sub>0.15</sub>FeO<sub>3-δ</sub>, respectively; and

Fig. 4 is a graph showing creep rates for an A-site rich metallic oxide of the present invention (Example No. 1 of Example 1) versus an A-site deficient composition (Example No. 2 of Example 1).

25

A preferred process for employing solid-state membranes comprising the metallic oxides of this invention is presented in US-A-6,077,323. This process will be illustrated by reference to the attached Fig. 1, which is a process flow diagram of this preferred process of this invention.

30

Oxygen-containing feed gas 1, preferably air, is pressurized in compressor 3 to a pressure in the range of 1 to 45 psig (0.07 to 3 barG), preferably less than 10 psig (0.7 barG). While air is the preferred oxygen-containing feed gas, other oxygen-containing gases can be utilized as an oxygen source for the process.

35 Pressurized oxygen containing gas 5 is preheated in heat exchanger 7 in heat trans-

fer zone 9, and preheated oxygen-containing gas 11 is heated further by direct combustion with fuel 14 in burner 15 to yield heated oxidant 17 typically containing 15-20 vol-% oxygen at a temperature above 500°C and preferably within  $\pm 200^\circ\text{C}$ , more preferably with  $\pm 110^\circ\text{C}$ , of the reactant feed temperature. Burner 15 can be  
5 any type of known, commercially available combustion device for promoting essentially complete combustion of fuel 14 in an excess oxygen environment.

Methane-containing reactant gas stream 19 is obtained at a pressure of 100-900 psig (7-62 barG), preferably 200-400 psig (14-28 barG), optionally by  
10 compression or pressure reduction (not shown) of a source gas. Methane-containing reactant gas stream 19 can be a methane-containing gas from a petroleum refinery, petrochemical plant, or other industrial source, or can be natural gas obtained from a pipeline or directly from the wellhead after appropriate pretreatment. Methane-containing reactant gas stream 19 preferably is natural gas  
15 with a typical composition of at least 80 vol% methane, less than 20 vol% ethane, less than 10 vol% propane, less than 5 vol% alkanes with more than 3 carbon atoms, less than 10 vol% carbon dioxide, less than 10 vol% nitrogen, and less than 100 ppmv total sulfur.

20 Methane-containing reactant gas stream 19 optionally is combined with a stream of hydrogen 21 and optionally heated in heat exchanger 23 in heat transfer zone 9 to a temperature up to 450°C. The resulting heated stream optionally is passed through desulfurization/hydrogenation reaction zone 25 containing a hydrogenation catalyst which typically comprises cobalt and molybdenum or nickel  
25 and molybdenum. In reaction zone 25, olefinic hydrocarbons are converted to paraffins and organic sulfur compounds are converted into hydrogen sulfide which is sorbed on a layer of zinc oxide in the reactor. Typically one reactor vessel containing hydrogenation catalyst is operated in series with two reactors filled with zinc oxide which are operated in parallel (not shown) with one onstream and one  
30 being regenerated.

Treated methane-containing gas 27 (which may contain residual hydrogen) is optionally combined with steam 29 and/or carbon dioxide 31, and the combined stream is heated to 500 to 800 °C in heat exchanger 33 in heat transfer zone 9 to  
35 yield heated reactant feed gas 35. Preferably steam is used, and the water to carbon molar ratio is between 0.5 to 5, the water to carbon molar ratio being defined

as the number of molecules of water in heated reactant feed gas 35 divided by the total number of carbon atoms present in the hydrocarbons and carbon monoxide in heated reactant feed gas 35.

5           Heated oxidant 17 and heated reactant feed gas 35 are introduced into mixed conducting membrane reaction zone 37. Mixed conducting membrane reaction zone 37 is shown schematically having an oxidant zone 43 separated from reactant zone 45 by mixed conducting membrane 47 comprising a dense layer formed from a metallic oxide according to Formula (I), and is presented in this simplified format for  
10           the following description of the reactor operation. Oxidant zone 43 represents a reactor volume through which the oxidant gas flows and contacts the oxidant side surface of the membrane 47. Dioxygen is ionized at this surface to form oxygen ions and the oxygen ions permeate the membrane 47 to the reactant side surface of the membrane. The surface of membrane 47 in oxidizing side 43 optionally can be  
15           coated with catalytic material to promote the transfer of oxygen into the membrane as discussed above.

          Reactant zone 45 represents a reactor volume through which the methane-containing reactant gas flows and reacts with oxygen which has permeated through  
20           the membrane 47. A number of known chemical reactions occur in reactant zone 45 among the several chemical species present including oxygen, hydrogen, water, carbon monoxide, carbon dioxide, methane, and possibly elemental carbon. For a further detailed discussion of these reactions and their implications on the process, the reader is referred to US-A 6,077,323.

25           The presence of hydrogen or steam in the feed is beneficial for preventing carbon deposition. A mixture of natural gas and steam would typically be limited to a preheat temperature of about 550°C. A mixture containing methane, steam, and hydrogen, but no hydrocarbons heavier than methane, could be heated to higher  
30           temperatures above 650°C depending on the relative concentrations of the components. Once the reactant gas enters reaction zone 37 and begins to react, the heavier hydrocarbons quickly disappear and a substantial amount of hydrogen is formed, so that cracking becomes progressively less likely in the successive zones of the reactor. The gradual permeation of oxygen through membrane 47 into the  
35           reactant is also beneficial in reducing the likelihood of carbon deposition.



The total gas pressure at any point in reactant zone 45 is 100-900 psig (7-62 barG), preferably 200-400 psig (14-28 barG), and a small pressure drop occurs from reactant zone inlet 41 to reactant zone outlet 49. The total gas pressure at any point in oxidant zone 43 is in the range of 1 to 45 psig (0.07-3 barG), preferably less than 10 psig (0.7 barG), and a small pressure drop occurs from oxidant zone inlet 39 to oxidant zone outlet 55. Thus the total pressure at any point in the reactant zone is greater than the total pressure at any point in the oxidant zone. The membrane of the invention withstands this pressure difference without cracking or deterioration of properties, especially of oxygen flux.

Hot syngas product 51 is withdrawn at outlet 49 at a temperature of greater than 1500°F (815°C) and cools in downstream equipment. Syngas product 51 contains hydrogen and carbon monoxide with a hydrogen to carbon monoxide molar ratio of 1 to 6. Oxygen-depleted oxidant 53 is withdrawn at outlet 55 at a temperature below that of product syngas 51.

Oxygen-depleted oxidant 53 is introduced into heat transfer zone 9 and exits therefrom as cooled flue gas 57. A major portion of the heat content of hot oxygen-depleted oxidant 53 is transferred via heat exchangers 7, 23, and 33 to heat process streams as earlier described. Heat transfer zone 9 can be similar to flue gas heat recovery systems used in conventional steam-methane reforming.

Hot syngas product 51 may be cooled rapidly to a temperature below 430°C against boiling water by indirect heat transfer in waste heat boiler 59 and is further cooled against other process streams (later defined) in one or more heat exchangers 61, 63, 65, and 67. Cooled syngas 69 passes into phase separator 71 from which condensed water 73 is withdrawn and combined with boiler feedwater makeup 75. The combined water stream is heated in heat exchanger 65 to yield preheated boiler feedwater 77 which is degasified and deaerated (not shown) for use in waste heat boiler 59. Alternatively, if the process makes excess water, a portion of condensate 73 is preheated in heat exchanger 65 and the remainder is discharged as wastewater (not shown). Depending on the end use of the syngas, cooled and dewatered syngas 79 optionally is treated in carbon dioxide removal system 81 using known methods to remove some or all of the carbon dioxide contained in the raw syngas product. Processed syngas 85 is compressed as required in compressor 87 to yield final syngas product 89.

Optionally, a portion of carbon dioxide 83 removed from the raw syngas is compressed in compressor 91 to provide carbon dioxide recycle 31 as earlier described. Optionally, a portion 93 of syngas 85 is separated using known methods such as pressure swing adsorption system 95 to recover hydrogen 21 which is used  
5 for hydrogenation of feed gas 19 as earlier described. Waste gas 97 from pressure swing adsorption system 95 can be combined with fuel 13 to provide fuel 14 in burner 15 for heating oxygen-containing feed gas 11.

In an alternative embodiment of the invention, steam 29 is not utilized and  
10 instead treated methane-containing gas 27 is directly saturated with water vapour before final heating and introduction into mixed conducting membrane reaction zone 37.

The present invention will be further illustrated by the following examples, which  
15 are given for illustrating purposes only and are not intended to limit the scope of protection.

#### EXAMPLE 1

##### PREPARATION OF MULTICOMPONENT METALLIC OXIDES

20

The compositions identified in Table 1 were prepared by known powder preparation techniques wherein the specified parts by weight of the respective metallic oxides were vibratory milled together for 72 hours. The mixture of metallic oxides was then fired in air at 1200°C for 10 hrs. Thereafter, the mixture was ground  
25 by vibratory milling for 72 hours yielding a powder. The powder was combined with a polyvinyl butyral binder, butyl benzyl phthalate plasticizer and an ethanol/toluene solvent to form a slip suitable for tape casting. The slip was cast into a tape and dried using conventional methods. Rectangular sections were cut from the tape using standard methods.

30

If necessary, several rectangular sections were laminated together to form a solid-state membrane having sufficient thickness. The solid-state membrane was fired in air to remove the plasticizer, binder and solvent followed by sintering at  
35 1400°C for 2 hours to produce a solid-state membrane having a thickness of about 2 millimeters consisting of a dense mixed conducting multicomponent metallic oxide layer.

TABLE 1

Ex No.	Composition	Ratio A/B	La <sub>2</sub> O <sub>3</sub> (pbw)	CaCO <sub>3</sub> (pbw)	SrCO <sub>3</sub> (pbw)	Fe <sub>2</sub> O <sub>3</sub> (pbw)	Al <sub>2</sub> O <sub>3</sub> (pbw)
1	(La <sub>0.85</sub> Ca <sub>0.15</sub> ) <sub>1.01</sub> FeO <sub>3-δ</sub>	1.01	0.5955	0.0646	-	0.3400	-
2*	(La <sub>0.85</sub> Ca <sub>0.15</sub> ) <sub>0.98</sub> FeO <sub>3-δ</sub>	0.98	0.5893	0.0639	-	0.3468	-
3*	(La <sub>0.85</sub> Sr <sub>0.15</sub> ) <sub>1.01</sub> FeO <sub>3-δ</sub>	1.01	0.5777	-	0.0924	0.3299	-
4*	(La <sub>0.75</sub> Ca <sub>0.25</sub> ) <sub>0.98</sub> FeO <sub>3-δ</sub>	0.98	0.5343	0.1094	-	0.3563	-
5*	(La <sub>0.15</sub> Sr <sub>0.85</sub> ) <sub>1.0</sub> Fe <sub>0.7</sub> Al <sub>0.3</sub> O <sub>3-δ</sub>	1.0	0.1105	-	0.5675	0.2528	0.0692
6	(La <sub>0.8</sub> Ca <sub>0.2</sub> ) <sub>1.01</sub> FeO <sub>3-δ</sub>	1.01	0.5681	0.0873	-	0.3446	-
7	(La <sub>0.9</sub> Ca <sub>0.1</sub> ) <sub>1.01</sub> FeO <sub>3-δ</sub>	1.01	0.6221	0.0425	-	0.3354	-
8	(La <sub>0.95</sub> Ca <sub>0.05</sub> ) <sub>1.01</sub> FeO <sub>3-δ</sub>	1.01	0.6480	0.0210	-	0.3310	-

\* not metallic oxide of the invention

- 5 Solid-state membranes consisting of a dense layer formed from the respective metallic oxides presented in Table 1 were tested for oxidant flux performance over time by applying an oxygen partial pressure gradient. Oxygen flux performance for Example Nos. 1 and 6 is shown in Fig. 2. These oxygen fluxes meet the requirements of commercial syngas production and are sufficiently stable with time.

10

#### EXAMPLE 2

#### EXPANSION TESTING OF SOLID-STATE MEMBRANES

- 15 Solid-state membrane formed from the metallic oxides identified in Ex. Nos. 1, 4, and 5 of Table 1 were prepared according to Example 1. Such solid-state membranes were tested under the following specified conditions for their expansion upon changing P<sub>O<sub>2</sub></sub>. The respective solid-state membranes were placed in a dilatometer and heated to 750°C / 950°C in pure oxygen. After equilibrating at 750°C, the atmosphere inside the dilatometer was switched to a H<sub>2</sub>/H<sub>2</sub>O/N<sub>2</sub> mixture having oxygen partial pressure given
- 20 in Table 2. The dilatometer recorded the change in length of the solid-state membranes with the change in oxygen partial pressure.

The results of the test are shown in Table 2 which demonstrates that a representative composition of the present invention (La<sub>0.85</sub>Ca<sub>0.15</sub>)<sub>1.01</sub>FeO<sub>3-δ</sub> shows

much lower expansion under pressure than the prior art compositions

$(\text{La}_{0.75}\text{Ca}_{0.25})_{0.98}\text{FeO}_{3-\delta}$  and  $(\text{La}_{0.15}\text{Sr}_{0.85})_{1.0}\text{Fe}_{0.7}\text{Al}_{0.3}\text{O}_{3-\delta}$ .

One of ordinary skill in the art will recognize that under typical syngas  
 5 process conditions, the solid-state membrane will be subjected to mechanical  
 stresses arising from the pressure gradient on opposite sides of the solid-state  
 membrane. Solid-state membranes comprising a dense layer formed from a  
 composition having a low expansion coefficient will be subjected to less mechanical  
 stress than solid-state membranes formed from a composition having a relatively  
 10 higher expansion coefficient. Thus, solid-state membranes formed from the metallic  
 oxides according to this invention will provide improved reliability and service life.

TABLE 2

Ex. No.	Composition	Final $\text{P}_{\text{O}_2}$ (atm) (Pa)	Expansion (ppm)
1	$(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{3-\delta}$	$5.5 \times 10^{-21}$ ( $5.5 \times 10^{-16}$ )	1820
4*	$(\text{La}_{0.75}\text{Ca}_{0.25})_{0.98}\text{FeO}_{3-\delta}$	$5.5 \times 10^{-21}$ ( $5.5 \times 10^{-16}$ )	4060
5*	$(\text{La}_{0.15}\text{Sr}_{0.85})_{1.0}\text{Fe}_{0.7}\text{Al}_{0.3}\text{O}_{3-\delta}$	$2.8 \times 10^{-13}$ ( $2.8 \times 10^{-8}$ )	3918

15 \* not metallic oxide of the invention

### EXAMPLE 3

#### STABILITY OF MULTICOMPONENT METALLIC OXIDES

20 To evidence improved stability afforded by the solid-state membranes of this  
 invention, the equilibrium conditions for the reaction of  $(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{2.925}$  and  
 $(\text{La}_{0.85}\text{Sr}_{0.15})_{1.01}\text{FeO}_{2.925}$  with  $\text{CO}_2$  to form  $\text{LaFeO}_3$ , alkaline earth carbonate, Fe and  
 oxygen at  $750^\circ\text{C}$  are shown in Figure 3. Fig. 3 shows equilibrium concentrations of  
 $\text{CO}_2$  and  $\text{O}_2$  for the reactions:

25  $(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{2.95} + 0.15\text{CO}_2 \rightleftharpoons 0.15\text{CaCO}_3 + 0.85\text{LaFeO}_3 + 0.004\text{La}_2\text{O}_3 +$   
 $0.15\text{Fe} + \frac{3}{4}(0.15)\text{O}_2$  and  
 $(\text{La}_{0.85}\text{Sr}_{0.15})_{1.01}\text{FeO}_{2.95} + 0.15\text{CO}_2 \rightleftharpoons 0.15\text{SrCO}_3 + 0.85\text{LaFeO}_3 + 0.004\text{La}_2\text{O}_3 + 0.15\text{Fe}$   
 $+ \frac{3}{4}(0.15)\text{O}_2$ .

These numbers were calculated using the FACT (*Facility for the Analysis of Chemical Thermodynamics*) integrated thermochemical database for the free energy of formation of the reactants and products. FACT was established as a joint research project between McGill University and Ecole Polytechnique de Montreal.

- 5 The free energy of formation of  $(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{2.925}$  and  $(\text{La}_{0.85}\text{Sr}_{0.15})_{1.01}\text{FeO}_{2.925}$  were calculated assuming ideal solution of  $\text{LaFeO}_3$  with  $\text{CaFeO}_{2.5}$  or  $\text{SrFeO}_{2.5}$  respectively. At a given oxygen partial pressure,  $(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{2.925}$  can be exposed to a much higher partial pressure of  $\text{CO}_2$  than  $(\text{La}_{0.85}\text{Sr}_{0.15})_{1.01}\text{FeO}_{2.925}$  before it will react with  $\text{CO}_2$ . Therefore,  $(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{2.925}$  is more stable than
- 10  $(\text{La}_{0.85}\text{Sr}_{0.15})_{1.01}\text{FeO}_{2.925}$ .  $(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{2.925}$  is more stable in applications requiring exposure to high partial pressures of  $\text{CO}_2$  at low oxygen partial pressures such as found in the reactant zone of a membrane reactor for syngas production, as described above.

15

#### EXAMPLE 4

##### CREEP MEASUREMENT OF SOLID-STATE MEMBRANES

- The solid-state membranes consisting of a dense layer of the respective compositions of Ex. Nos. 1 and 2 according to Table 1, were prepared according to
- 20 the procedure of Example 1. Each solid-state membrane was individually placed in an apparatus using a four point bend arrangement to measure creep rate. A description of a suitable technique for this measurement can be found in Chapter 5 of "Mechanical Testing Methodology for Ceramic Design and Reliability" edited by David C. Cramer and David W. Richerson published by Marcel Dekker, Inc., 1998.
- 25 Each individual solid-state membrane was heated to 950°C and stresses of 1000-6000 psi (1-6 ksi) (6.9-41.4 MPa) were applied thereto.

- The deflection of the respective solid-state membranes was measured as a function of time, from which the strain rate was calculated. The creep rates are
- 30 presented in Fig. 4 for the solid-state membranes formed from Ex. Nos. 1 ( $A/B=1.01$ ) and 2 ( $A/B=0.98$ ). The creep rates for the solid-state membrane formed from Composition No. 1 are nearly an order of magnitude lower than the creep rates of the solid-state membrane formed from Composition No. 2. Therefore compositions with an  $A/B$  ratio  $> 1.0$  have significantly lower creep rates than membranes with an
- 35  $A/B$  ratio  $< 1.0$ .

**EXAMPLE 5**  
**SYNGAS PRODUCTION**

A tubular solid-state membrane consisting of a dense layer of the composition  
5 (La<sub>0.85</sub>Ca<sub>0.15</sub>)<sub>1.01</sub>FeO<sub>3.8</sub>, was fabricated by consolidating a mixture of the ceramic oxide powder, polyvinylbutyral binder and butyl benzyl phthalate plasticizer, by isostatic pressing in a shaped tool. The pressed tube was then fired in air in a controlled manner to remove the plasticizer, followed by sintering at 1400°C for 2 hours to produce a  
10 dense mixed conducting multicomponent metallic oxide membrane in form of a tube having a thickness of about 0.95 mm.

After firing, the exterior surface of the tube was machined to final dimensions and coated externally with an approximately 100 µm thick porous layer of a surface reforming catalyst consisting of 30% by weight Ni dispersed on (La<sub>0.75</sub>Ca<sub>0.25</sub>)<sub>1.01</sub>FeO<sub>3.8</sub>.  
15 The tube was also coated internally with an approximately 10 µm thick porous layer of an oxygen reduction catalyst, La<sub>0.49</sub>Sr<sub>0.49</sub>CoO<sub>3.8</sub>. The catalyst layers were fired onto the tube prior to testing in order to partially consolidate the catalyst particles and in order to remove any organic components.

20 The resulting tubular solid-state membrane was closed on one end and had an average wall thickness of 0.95 mm and an mean active surface area of about 14.5 cm<sup>2</sup>. The open end of the tubular membrane was sealed into a Haynes 230 alloy tube with a ceramic/metal compression seal, and placed in a heated tubular reactor vessel fitted with an alumina liner. An alumina air feed tube was inserted into the interior of the  
25 sample such that the pre-reformed natural gas feed mixture and the air feed flowed in the reactor vessel in a largely co-axial manner.

A pre-reformed natural gas mixture (composition given in Table 3) was supplied to the outside of the membrane tube at a pressure of 250 psia (1.7 MPa) and a flow rate  
30 of approximately 730 standard cm<sup>3</sup>/min, while air at atmospheric pressure was contacted with the inside of the tubular solid-state membrane at a flow rate of approximately 750 standard cm<sup>3</sup>/min. The average temperature of the tubular solid-state membrane was 810°C.

**TABLE 3**  
**PRE-REFORMED NATURAL GAS FEED COMPOSITION**

Gas Component	Mole %
CH <sub>4</sub>	4.80
CO	16.85
CO <sub>2</sub>	10.68
H <sub>2</sub>	41.92
H <sub>2</sub> O	25.75
Total	100.00

5 Under these conditions, oxygen was transported through the tubular solid-state membrane to partially oxidize the pre-reformed natural gas mixture. A typical product gas composition is given in Table 4. The oxygen flux through the tubular solid-state membrane was calculated by performing a mass balance on both the natural gas mixture feed and oxidized product stream, and on the air feed and oxygen depleted exhaust stream. Over a period of fifteen days continuous operation, the tubular solid-state membrane exhibited a stable flux of oxygen from the air to the pre-reformed natural gas mixture of approximately 2.0 standard cm<sup>3</sup> oxygen/min/cm<sup>2</sup> of active membrane surface.

**TABLE 4**  
**PARTIALLY OXIDIZED PRODUCT GAS COMPOSITION**

Gas Component	Mole %
CH <sub>4</sub>	4.05
CO	14.21
CO <sub>2</sub>	12.58
H <sub>2</sub>	37.05
H <sub>2</sub> O	32.11
Total	100.00

20 Those skilled in the art will appreciate that the claimed multicomponent metallic oxides are particularly suited toward use in fabricating solid-state membranes suitable for producing syngas.

## Claims:

1. A metallic oxide represented by Formula (I)  
$$(\text{Ln}_x\text{Ca}_{1-x})_y\text{FeO}_{3-\delta} \quad (\text{I})$$
- 5 wherein  
Ln is La or a mixture of lanthanides comprising La, and wherein  
 $1.0 > x > 0.5$   
 $1.1 \geq y > 1.0$  and  
 $\delta$  is a number which renders the metallic oxide charge neutral.
- 10 2. A metallic oxide of Claim 1 wherein  
 $0.98 > x > 0.75$  and  
 $1.05 \geq y \geq 1.01$ .
- 15 3. A metallic oxide of Claim 1 selected from  $(\text{La}_{0.85}\text{Ca}_{0.15})_{1.01}\text{FeO}_{3-\delta}$ ;  
 $(\text{La}_{0.8}\text{Ca}_{0.2})_{1.01}\text{FeO}_{3-\delta}$ ;  $(\text{La}_{0.9}\text{Ca}_{0.1})_{1.01}\text{FeO}_{3-\delta}$ ; and  $(\text{La}_{0.95}\text{Ca}_{0.05})_{1.01}\text{FeO}_{3-\delta}$ .
- 20 4. A solid-state membrane which comprises a dense layer formed from a metallic  
oxide as defined in any one of Claims 1 to 3.
5. A solid-state membrane of Claim 4 which further comprises a porous  
mixed conducting multicomponent metallic oxide layer contiguous to the said dense  
layer.
- 25 6. A solid-state membrane of Claim 4 or Claim 5 which further comprises a  
catalyst on a first side thereof, a catalyst on a second side thereof or a catalyst on the  
first side and the second side thereof.
- 30 7. A solid-state membrane of Claim 6 having on the first side thereof a  
reforming or partial oxidation catalyst for synthesis gas production.
8. A solid-state membrane of Claim 6 or Claim 7 said catalyst on the first  
side thereof comprises a metal or an oxide of a metal selected from Groups 5, 6, 7, 8, 9,  
10, 11 of the Periodic Table of the Elements according to the International Union of Pure  
35 and Applied Chemistry.



9. A solid-state membrane of Claim 8 wherein the catalyst comprises a metal or an oxide of a metal selected from nickel, cobalt, iron, platinum, palladium, and silver.

5 10. A solid-state membrane of any one of Claims 6 to 9 having on the second side thereof an oxygen reduction catalyst.

11. A solid-state membrane of Claim 10 wherein said catalyst on the second side thereof comprises a metal or an oxide of a metal selected from the Groups 2, 5, 6,  
10 7, 8, 9, 10, 11 and 15 and the F block lanthanides of the Periodic Table of the Elements according to the International Union of Pure and Applied Chemistry.

12. A solid-state membrane of Claim 11 wherein the said metal or the oxide of the metal is selected from platinum, palladium, ruthenium, gold, silver, bismuth,  
15 barium, vanadium, molybdenum, cerium, praseodymium, cobalt, rhodium and manganese.

13. A solid-state membrane of any one of Claims 6 to 9 having on the second side thereof a multicomponent metallic oxide.  
20

14. A process for producing a synthesis gas product comprising hydrogen and carbon monoxide which comprises the steps of:

i) providing a reaction zone having an oxidant side and a reactant side  
25 which are separated by a solid-state membrane as defined in any one of Claims 4 to 13,

ii) contacting a heated oxygen-containing feed gas with the oxidant side of the said membrane at an oxidant feed temperature and an oxidant gas feed pressure;

30 iii) contacting a heated methane-containing reactant gas with the reactant side of the said membrane at a reactant gas feed temperature and a reactant gas feed pressure;

whereby oxygen from the oxidant side of the reaction zone permeates through  
35 the solid-state membrane to the reactant side of the reaction zone and reacts with the methane-containing reactant gas to form the synthesis gas product;

iv) withdrawing the synthesis gas product from the reactant side of the reaction zone; and

5 v) withdrawing an oxygen depleted gas stream from the oxidant side of the reaction zone.

15. A process of Claim 14 further comprising the steps of:

10 vi) contacting a heated gaseous stream comprising steam and one or more hydrocarbons with at least one catalyst which promotes steam reforming of hydrocarbons to form a partially reformed intermediate gas comprising at least methane, hydrogen and carbon oxides; and

15 vii) introducing the partially reformed intermediate gas into the reactant side of the reaction zone of step iii).

20 16. A process of Claim 14 or Claim 15 wherein the reactant feed gas temperature is between 510°C (950°F) and 760°C (1400°F) and the synthesis gas product outlet temperature from the reactant side is greater than 815°C (1500°F).

17. A process of any one of Claims 14 to 16 wherein the oxidant gas feed pressure is between 0.07 barG (1 psig) and 3 barG (45 psig).

25 18. A process of any one of Claims 14 to 17 wherein the reactant gas feed pressure is between 7 barG (100 psig) and 62 barG (900 psig).

19. A process of any one of Claims 14 to 18 wherein the oxidant gas feed temperature is up to 95°C (200°F) greater than the reactant gas feed temperature.

30 20. A process of any one of Claims 14 to 19 wherein the oxidant gas feed temperature is less than the oxygen-depleted oxidant gas outlet temperature.

35 21. A process of any one of Claims 14 to 20 wherein the oxygen containing feed gas in step (ii) is heated by direct combustion with a fuel in a direct-fired combustor to produce a hot, pressurized combustion product to provide the heated oxygen-containing feed gas.

22. A process of any one of Claims 14 to 21 wherein the methane-containing reactant gas further comprises one or more components selected from water, carbon dioxide and hydrogen.

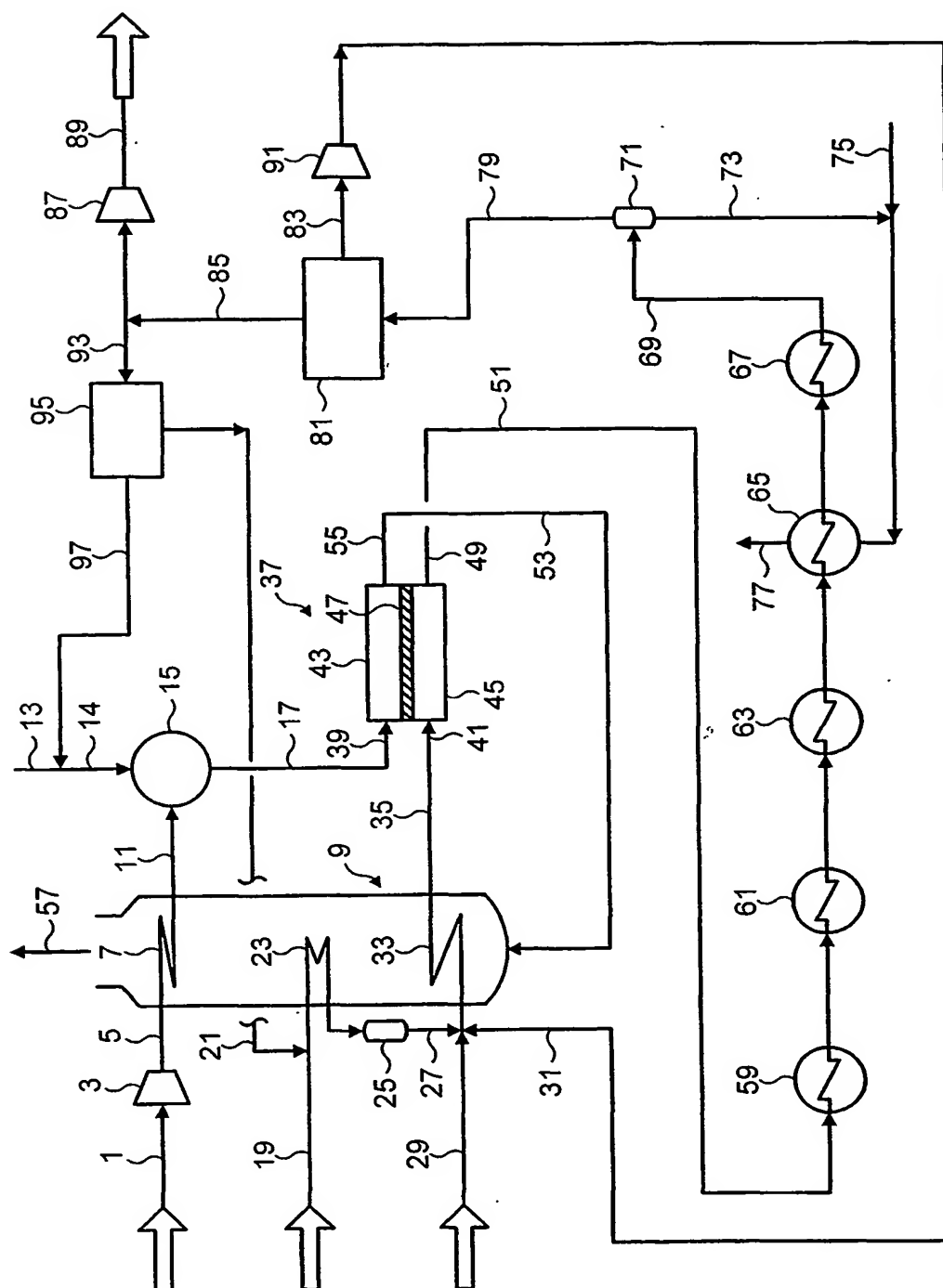


FIG. 1

2 / 4

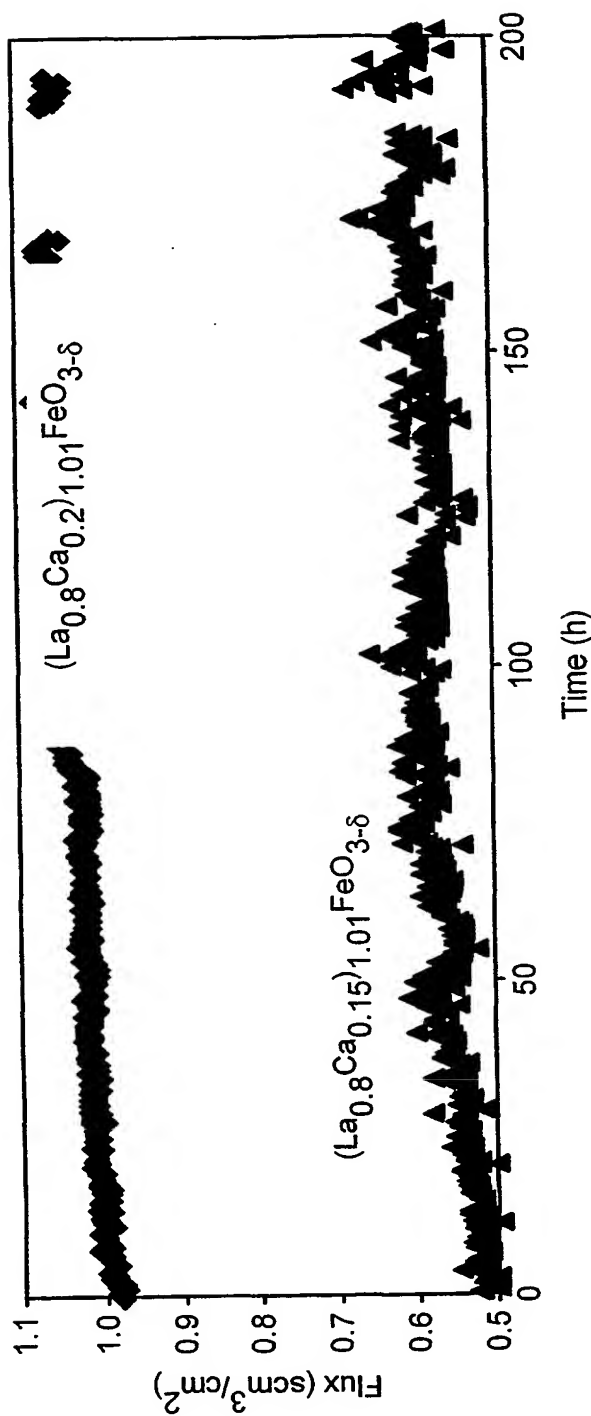


FIG. 2

3 / 4

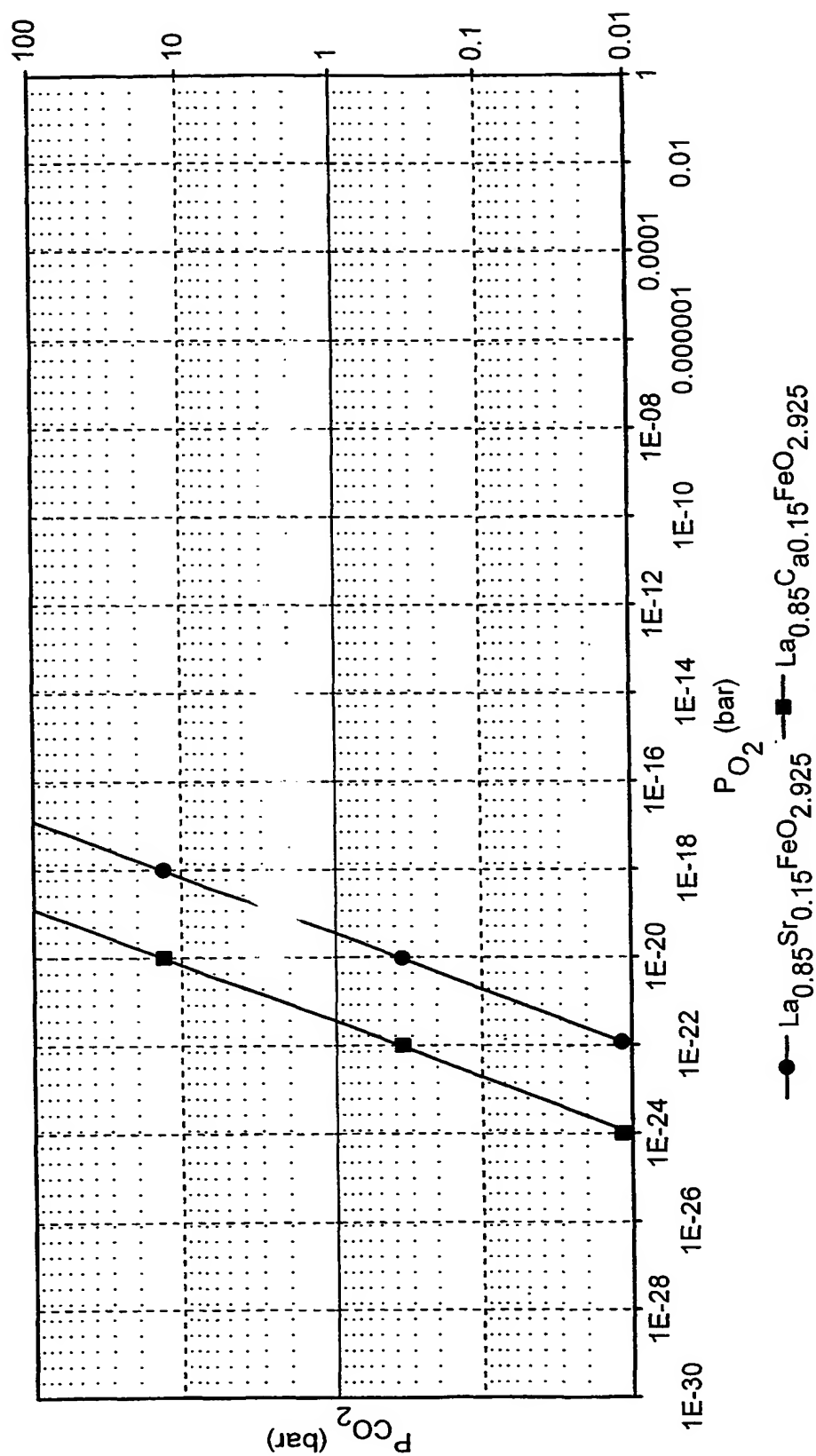


FIG. 3

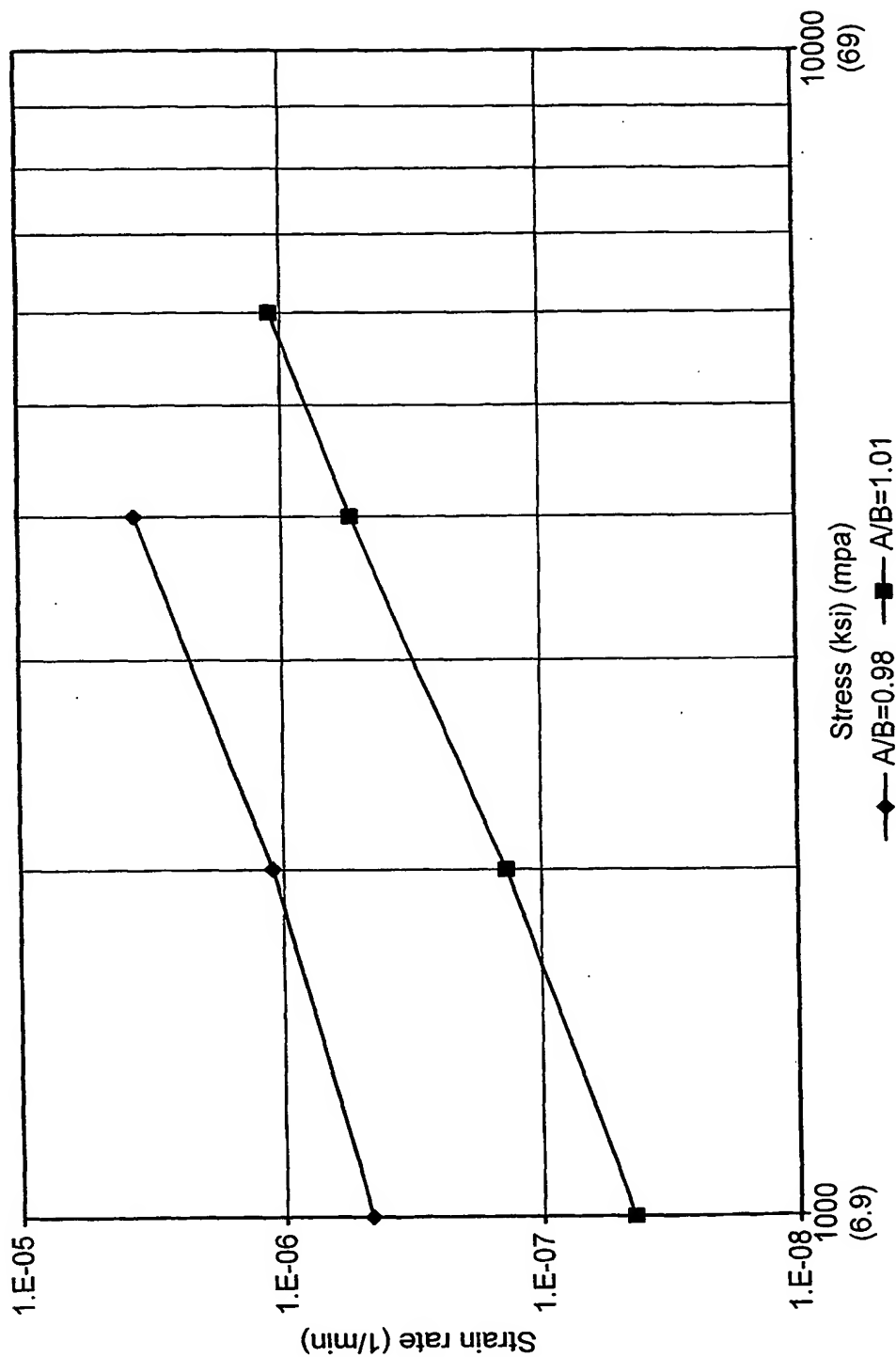


FIG. 4

## INTERNATIONAL SEARCH REPORT

Inter <sup>national</sup> Application No  
PCT/EP 01/09704

## A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 B01D71/02 B01D53/22 B01J12/00 C01B3/34 C04B35/26

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 B01D B01J C01B C04B

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

WPI Data, CHEM ABS Data

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 5 972 296 A (HERAEUS ELECTRO-NITE INTERNATIONAL, N.V.) 26 October 1999 (1999-10-26)	1-3
Y	column 2, line 21 - line 28; claims 1-11	4-22
X	L.A. ISUPOVA ET AL.: KINETICS AND CATALYSIS, vol. 41, no. 2, February 2000 (2000-02), pages 287-291, XP001037684 the whole document	1-3
Y	US 6 033 632 A (ELTRON RESEARCH, INC.) 7 March 2000 (2000-03-07) cited in the application the whole document	4-22
	--- -/-	

☒ Further documents are listed in the continuation of box C.☒ Patent family members are listed in annex.

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Date of the actual completion of the international search

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Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2  
NL - 2280 HV Rijswijk  
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,  
Fax (+31-70) 340-3016

Authorized officer

Luethe, H



# INTERNATIONAL SEARCH REPORT

International Application No  
PCT/EP 01/09704

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